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11. In your own words, what is the difference between a sublevel and an orbital? Orbitals are contained within sublevels. A sublevel is a grouping of orbitals. 12. How many electrons can fit in each of the following energy levels: 1st energy level = 2 because only 2 can fit in an s sublevel and the first energy level only has an s sublevel.

Information: Energy Levels and Sublevels

141 ChemQuest 46 Name: _____ Page 143 and 144: 143 9. Consider your answers to que. Page 145 and 146: 145 Critical Thinking Questions 5. Page 147 and 148: 147 8. If a substance is a Bronsted. Page 149 and 150: 149 6. Use the information given an. Page 151 and 152: 151 ChemQuest 50 Name: _____ Page 153 and 154: 153 11. Find the pH of a 0.85 M sol

34 ChemQuest 11 Name:

Chemquest 11 Answers 34 ChemQuest 11 Name: _____ Date: _____ Hour: _____ Information: Energy Levels and Sublevels As you know, in his solar system model Bohr proposed that electrons are located in energy levels. The current model of the atom isn't as simple as that, however. Anatomy of a ChemQuest

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11. The energy difference between the 1st and 2nd energy levels in a certain atom is 5.01×10^{-19} J. What frequency of light is necessary to excite an electron in the 1st energy level? 12. a) What is the frequency of light given off by an electron that loses 4.05×10^{-19} J of energy as it moves from the 2nd to the 1st energy level?

Critical Thinking Questions: Bohr's reasoning

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11. If you have a container with hydrogen gas and oxygen gas in it do you have water? Why or why not? You do not have water because the hydrogen and oxygen atoms are simply mixed together; they are not bonded together. 12. Give an example of something that is an element. Your example should not already be on this sheet.

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