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Chapter 13 States Of Matter

Chapter 13 States of Matter
137 SECTION
13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of

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the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386)
1.

Name Date Class **STATES OF MATTER** **13**

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all matter consists of tiny particles that are constantly in motion
What are the three assumptions of the kinetic theory as it applies to gases? -The particles in a gas are considered to be small, hard spheres with an insignificant volume.
-The motion of the

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particles in a gas are rapid, constant, and random.

Chapter 13: States of Matter Flashcards | Quizlet

There are three states of matter that we will learn about in this chapter. (If you want to learn about more states of matter, I can refer you to somebody.) Those three states are solid, liquid, and gas. These

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three states are quite different. The main difference is in their particles.

Chapter 13: States of Matter -

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No attractive or repulsive forces exist between the particles.

3 Chapter 13 States Of Matter Worksheet Title:

Chapter 13 States of Matter 1 Chapter 13

States of Matter 2

Kinetic Theory as

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Applied to Gases
Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume.

Chapter 13 States Of Matter Worksheet

→ Look at the text on page 315 for the answer. You are already familiar with the three common states of matter: solid,

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liquid, and gas. Solid objects litter the room around you. For example, you can easily recognize the shape of your desk; you know that your backpack cannot hold seven textbooks.

Chapter 13: States of Matter

Title: Chapter 13 States of Matter 1 Chapter 13 States of Matter 2 Kinetic Theory as Applied to Gases

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Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume. Between particles in a gas there is empty space. No attractive or repulsive forces exist between the particles. 3

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Matter. STUDY.

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bill_brossman. Terms in

this set (26) Kinetic

Energy. The energy an

object has because of

its motion. Kinetic

Theory. A theory that

explains the states of

matter, based on the

concept that all matter

consists of tiny

particles that are ...

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Chapter 13 - States of Matter - 13.3 The

Nature of Solids - 13.3

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Lesson Check - Page

434: 24 Answer Solids

have a definite shape
and a definite volume
because of the
arrangement of the
particles.

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The kinetic theory
states that the tiny
particles in all forms of
matter are in constant
motion! Section 13.1

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The Nature of Gases

Three basic
assumptions of the
kinetic theory as it
applies to...

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Chapter 13 "States of
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- OBJECTIVES:
- Describe the assumptions of the "kinetic theory" as it applies to gases. Feb 109:15 AM

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- OBJECTIVES:
- Interpret gas pressure in terms of kinetic theory. Feb 109:15 AM
- OBJECTIVES: • Define the relationship between Kelvin temperature

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ISBN-10: 0132525763,
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Chemistry (12th Edition) Chapter 13 - States of Matter ...

all matter consists of
tiny particles that are
in constant motion:

normal boiling point:

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the boiling point of a
Liquid at a pressure of
101.3 kPa: phase
diagram: graph
representing the
relationships among
the solid, liquid, and
vapor states of a
substance in a sealed
container: standard
atmosphere: pressure
required to support
760 mm of mercury ...

Quia - Chapter 13 "States of Matter"

Chapter 13 - States of
Page 21/26

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- Page 443: 54 Answer

Temperature stays constant because the energy is needed to break the intermolecular forces and changing the phase.

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434: 18 Answer Atoms, Ions, or molecules are packed tightly together in an orderly arrangement.

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